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**QUESTION 1**

What is the conjugate base of sulfuric acid?

- A. H_2SO_3
- B. HSO_4^-
- C. SO_4^{2-}
- D. H_3SO_4^+

A. Option A

B. Option B

C. Option C

D. Option D

Correct Answer: B

QUESTION 2

Which of the following is not secreted by the hypothalamus?

- A. glucocorticoids
- B. GnRH
- C. Dopamine
- D. GHIH

Correct Answer: A

Glucocorticoids are steroids synthesized and secreted by the adrenal cortex not the hypothalamus.

QUESTION 3

Blood flows from the mitral valve to the:

- A. left atrium
- B. aorta



C. right ventricle

D. left ventricle

Correct Answer: D

The left atrium connects to the left ventricle by way of the mitral valve. The left ventricle connects to the aorta by way of the aortic valve. Blood flows from the pulmonary veins into the left atrium and through the mitral valve (bicuspid valve), to the left ventricle then out through the aorta.

QUESTION 4

Which of the following atoms has the highest electronegativity?

A. Cl

B. Br

C. N

D. F

Correct Answer: D

In the covalent bond, two atoms are joined by sharing electrons. Both nuclei are held by the same electron cloud. However, in most cases the two nuclei do not share the electrons equally. This happens when one atom has more electron withdrawing power than the other atom. At this time the electron cloud is denser on one atom. This result in one end of the bond being relatively negative and the other end being relatively positive. Such a bond is said to be a polar bond or to possess polarity. The bond possesses polarity when joins atoms have different tendency to attract electrons. This property of the atom is called electronegativity. Out of the given choices, fluorine (F) possess the highest electronegativity. $F > O > Cl, N > Br > C, H$

QUESTION 5

200 mL of an ideal gas is placed in a piston and is held at a pressure of 500 torr. If the temperature is held constant and the pressure is increased to 650 torr, what is the new volume of the gas?

A. 75 mL

B. 154 mL

C. 220 mL

D. 300 mL

Correct Answer: B

When the pressure of a gas increases at constant temperature, the volume decreases. Therefore, before calculation, choices [220 mL] and [300 mL] can be eliminated. Since PV / nT is constant, and the temperature and number of moles of gas are kept constant, we can derive the relationship, $P_1V_1 = P_2V_2$. Rearranging this equation to find the final volume of gas, we get: $V_2 = P_1V_1/P_2$. By substituting numbers into this equation, we get: $V_2 = (500 \times 200) / 650$, or 154 mL, choice B. Another way to arrive at the correct answer is to notice that the original pressure (500 mL) divided by the new pressure (650 mL) is approximately $3/4$, so the new volume is around $3/4$ of 200 mL or 150 mL.



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